

- Substitution of phosphine for CO ligand in $\text{HW}_2(\text{CO})_9(\text{NO})$. The structures of $\text{HW}_2(\text{CO})_8(\text{NO})(\eta^1-(\eta^5\text{-C}_5\text{H}_4\text{PPh}_2)_2\text{Fe})$, $\text{HW}_2(\text{CO})_7(\text{NO})(\text{Ph}_2\text{PH})_2$, $\text{HW}_2(\text{CO})_7(\text{NO})(\eta^2\text{-Ph}_2\text{PCH}_2\text{PPh}_2)$, and $[\text{HW}_2(\text{CO})_8(\text{NO})]_2(\mu\text{-Ph}_2\text{PCH}_2\text{CH}_2\text{PPh}_2)$ (J. Organomet. Chem., 388 (1990) 151–167)

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The last two lines of page 161 and the first line of page 162 should be corrected as follows:

$\text{HW}_2(\text{CO})_9(\text{NO})$ (Fig. 6), the phosphorus atoms prefer the exo positions 8 (in complexes 4, 5, 7, 9, and $\text{HW}_2(\text{CO})_8(\text{NO})(\text{P}(\text{OCH}_3)_3)$ [66]) and/or 9 (in complexes 5 and 7) to the positions 6 and 7. Such a substitution is likely to minimize steric

On page 165, Table 4, the ^1H NMR data for complex 6 (2nd line) should be corrected as follows:

W-H (-9.27, d, 1 H, $^2J(\text{P-H}) = 11.4$,

- Spectroelectrochemistry of aromatic ligands and their derivatives. III. Binuclear transition metal complexes of Cu^{I} , Mo^0 , and Re^{I} with 2,2'-bipyrimidine (J. Organomet. Chem., 411 (1991) 207–213)

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Table 1 that appears on p. 209 should be replaced by the following:

Table 1

Reduction potentials of 2,2'-bipyrimidine and its dinuclear complexes in DMF ^a

Compounds	1st reduction	2nd reduction	Difference
bpym ^b	-2.102(0.063) ^c	-2.619(irr) ^d	> 0.51
I	-0.997(0.064)	-1.641(0.079)	0.644
II	-0.680(0.072)	-1.355(0.092) ^e	0.675
III ^f	-1.090(0.064)	-1.730(0.079)	0.640

^a Data by cyclic voltammetry (50–200 mV s⁻¹, PAR 173/175; *iR* compensation), V vs. ferrocene/ferrocenium⁺ in stated solvent at 25°C. Measurements taken vs. Ag/0.01 M AgNO₃-0.09 M n-tetra-butylammonium tetrafluoroborate (TBABF₄) in stated solvent, but referred to ferrocene/ferrocenium⁺/0.1 M TBABF₄ in solvent/cell combination as used. ^b After ref. 12. ^c $E_{\text{pa}} - E_{\text{pc}}$ (V) in parentheses. ^d Denotes a chemically irreversible reduction process (cathodic peak potentials given at 200 mV s⁻¹ scan rate). ^e Incompletely chemically reversible, but anodic return wave detected. ^f With 0.01 M triphenylphosphine.